Biology AQA Specification

		4.1.1.1 Eukaryotes and prokaryotes
		4.1.1.2 Animal and plant cells
		4.1.1.3 Cell specialisation
	4.1.1 Cell Structure	4.1.1.4 Cell differentiation
		4.1.1.5 Microscopy
4.1 Cell biology		4.1.1.6 Culturing microorganisms (biology only)
		4.1.2.1 Chromosomes
	4.1.2 Cell division	4 1.2.2 Mitosis and the cell cycle
		4.1.2.3 Stem cells
		4.1.3.1 Diffusion
	4.1.3 Transport in cells	4.1.3.2 Osmosis
		4 1.3.3 Active transport
	4.2.1 Principles of organisation	
		4.2.2.1 The human digestive system
		4 2 2 2 The heart and blood vessels
		4.2.2.3 Blood
4.2 Organisation	4.2.2 Animal tissues, organs and organ	1.2.2.5 Brood
4.2 Organisation	systems	communicable disease
		4 2 2 5 Health issues
		4.2.2.6 The effect of lifestyle on some non-
		communicable diseases
		4.2.2.7 Cancer
	4.2.3 Plant tissues, organs and systems	4.2.3.1 Plant tissues
	4.2.5 Flant lissues, organs and systems	4 2 3 2 Plant organ system
		4 3 1 1 Communicable (infectious) diseases
		4 3 1 2 Viral diseases
		4 3 1 3 Bacterial diseases
	4 3 1 Communicable diseases	4 3 1 4 Fungal diseases
	4.0.1 Communicable discuses	4 3 1 5 Protist diseases
4.3 Infection and		4 3 1 6 Human defence systems
response		4.3.1.7 Vaccination
		4.3.1.9 Antibiotics and painkillers
		4.3.1.9 Discovery and development of drugs
		4.3.2.1 Producing monoglonal antibodies
	4.3.2 Monocional antibodies (biology	4.3.2.2 Uses of monoclonal antibodies
		4.3.3.1 Detection and identi action of plant
	4.2.2 Plant disease (hislamy only)	diseases
	4.5.5 Fiant disease (biology only)	4.3.3.2 Plant defence responses
		4.4.1.1 Photosynthetic reaction
	4.4.1 Photosynthesis	4.4.1.2 Rate of photosynthesis
4.4 Bioenergetics		4 4 1 3 Uses of glucose from photosynthesis
		4.4.2.1 Aerobic and anaerobic respiration
	4.4.2 Respiration	4 4 2 2 Response to exercise
		4 4 2 3 Metabolism
	4.5.1 Homeostasis	
		4.5.2.1 Structure and function
	4.5.2 The human nervous system	4.5.2.2 The brain (biology only)
		4.5.2.3 The eye (biology only)
		4.5.2.4 Control of body temperature (biology
		only)
		4.5.3.1 Human endocrine system
4.5 Homeostasis and	4.5.3 Hormonal coordination in humans	4.5.3.2 Control of blood glucose concentration
response		4.5.3.3 Maintaining water and nitrogen balance
		in the body (biology only)
		4.5.3.4 Hormones in human reproduction
		4.5.3.5 Contraception
		4.5.3.6 The use of hormones to treat infertility
		(HT only)
		4.5.3.7 Negative feedback (HT only)
	4.5.4 Plant hormones (biology only)	4.5.4.1 Control and coordination
		4.5.4.2 Use of plant hormones (HT only)
		4.6.1.1 Sexual and asexual reproduction
		4.6.1.2 Meiosis
		4.6.1.3 Advantages and disadvantages of sexual
		and asexual reproduction (biology only)
4.C. In heritenes	4.6.1 Reproduction	4.6.1.4 DNA and the genome

Key Triple Science Only HL only

4.0 IIIIEIIIaIICE.	1	
variation and		4.6.1.5 DNA structure (biology only)
evolution		4.6.1.6 Genetic inheritance
		4.6.1.7 Inherited disorders
		4.6.1.8 Sex determination
	4.6.2 Variation and evolution	4.6.2.1 Variation
		4.6.2.2 Evolution
		4.6.2.3 Selective breeding
		4.6.2.4 Genetic engineering
		4.6.2.5 Cloning (biology only)
		4.6.3.1 Theory of evolution (biology only)
		4.6.3.2 Speciation (biology only)
	4.6.3 The development of understanding of genetics and evolution	4.6.3.3 The understanding of genetics (biology only)
		4.6.3.4 Evidence for evolution
		4.6.3.5 Fossils
		4.6.3.6 Extinction
		4.6.3.7 Resistant bacteria
		4.6.4 Classi cation of living organisms
	4.7.1 Adaptations, interdependence and	4.7.1.1 Communities
	competition	4.7.1.2 Abiotic factors
		4.7.1.3 Biotic factors
		4.7.1.4 Adaptations
		4.7.2.1 Levels of organisation
	4.7.2 Organisation of an ecosystem	4.7.2.2 How materials are cycled
		4.7.2.3 Decomposition (biology only)
		4.7.2.4 Impact of environmental change
4.7 Ecology		(biology only) (HT only)
	4.7.3 Biodiversity and the effect of human interaction on ecosystems	4.7.3.1 Biodiversity
		4.7.3.2 Waste management
		4.7.3.3 Land use
		4.7.3.4 Deforestation
		4.7.3.5 Global warming
		4.7.3.6 Maintaining biodiversity
	4.7.4 Trophic levels in an ecosystem (biology only)	4.7.4.1 Trophic levels
		4.7.4.2 Pyramids of biomass
		4.7.4.3 Transfer of biomass
	4.7.5 Food production (biology only)	4.7.5.1 Factors affecting food security
		4.7.5.2 Farming techniques
		4.7.5.3 Sustainable sheries
		4.7.5.4 Role of biotechnology

Chemistry AQA Specification

Key	
Triple Science Only	
HL only	

		4.1.1.1 Atoms, elements and compounds
	4.1.1 A simple model of the atom, symbols,	4.1.1.2 Mixtures
	relative atomic mass, electronic charge and	4.1.1.3 The development of the model of the
	isotopes	atom (common content with physics)
		4.1.1.4 Relative electrical charges of subatomic
		particles
		4 1 1 5 Size and mass of atoms
		4.1.1.6 Paletiva atomia mass
4 1 Atomic structure and the		
periodic table		4.1.1.7 Electronic structure
		4.1.2.1 The periodic table
		4.1.2.2 Development of the periodic table
	4.1.2 The periodic table	4.1.2.3 Metals and non-metals
		4 1 2 4 Group 0
		4.1.2.5 Group 1
	4.1.3 Properties of transition metals (chemistry	4.1.2.5 Group 1
		4.1.2.6 Group /
		4.1.3.1 Comparison with Group 1 elements
	only)	4.1.3.2 Typical properties
		4.2.1.1 Chemical bonds
		4.2.1.2 Ionic bonding
		4 2 1 3 Ionic compounds
		4.2.1.4 Covalant handing
	4.2.1 Chemical bonds, ionic, covalent and metallic	4.2.1.5 Metallic bonding
		4.2.2.1 The three states of matter
		4.2.2.2 State symbols
4.2 Bonding, structure, and		4.2.2.3 Properties of ionic compounds
the properties of matter		4.2.2.4 Properties of small molecules
		4 2 2 5 Polymers
		4.2.2.6 Cient equalent atmostures
		4.2.2.7 Properties of metals and alloys
		4.2.2.8 Metals as conductors
		4.2.3.1 Diamond
		4.2.3.2 Graphite
	4.2.3 Structure and bonding of carbon	4.2.3.3 Graphene and fullerenes
	4.2.4 Dulk and surface nuoneuties of motton including	4.2.4.1 Sizes of particles and their properties
	4.2.4 Bulk and surface properties of matter including	4.2.4.2 Uses of particles and then properties
	nanoparticles (chemistry omy)	4.2.4.2 Uses of nanoparticles
		4.3.1.1 Conservation of mass and balanced
		chemical equations
	4.3.1 Chemical measurements, conservation of mass and the quantitative interpretation of chemical equations	4.3.1.2 Relative formula mass
		4.3.1.3 Mass changes when a reactant or product
		is a gas
		4 3 1 4 Chemical measurements
		1.5.1.1 Chemiear measurements
		4.3.2.1 Moles (HT only)
		4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT
		4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only)
	4.3.2 Use of amount of substance in relation to	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Lising moles to balance equations (HT
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only)
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only)
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Connectation of solutions
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only)	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 	4.3.2.1 Moles (HT only) 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only)
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only)	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom encomy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only)
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.1.1 Metal oxides 4.1.2 The reactivity series
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 	 4.3.2.1 Moles (HT only) 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.2 The reactivity series 4.4.1.2 The reactivity series
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.2.4 tom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.4 Oxidation and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only)
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only)
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.2 The reactivity series 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.4.1 Reactions of acids with metals
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.2.1 Reactions of acids with metals 4.4.2.1 Neutralisation of acids and salt metalis
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.3 Soluble salts
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.1.1 Metal oxides 4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.3 Soluble salts 4.4.2.4 The pH scale and neutralisation
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.2.2 The reactivity series 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.3 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only)
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids	 4.3.2.1 Moles (HT only) 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.1 Percentage yield 4.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.2 Neutralisation of acids with metals 4.2.2 Neutralisation of acids and salt production 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.2.6 Strong and weak acids (HT only)
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.2.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.1 Acxidation and reduction in terms of electrons (HT only) 4.4.2.1 Reactions of acids with metals 4.4.2.3 Soluble salts 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.2.6 Strong and weak acids (HT only)
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.2.1 Reactions of acids with metals 4.4.2.3 Soluble salts 4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.3.1 The process of electrolysis
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.1.1 Metal oxides 4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.2.1 Reactions of acids with metals 4.2.2 Neutralisation of acids and salt production 4.2.3 Soluble salts 4.2.4 The pH scale and neutralisation 4.2.5 Titrations (chemistry only) 4.3.1 The process of electrolysis 4.3.2 Electrolysis of molten ionic compounds
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 4.4.3 Electrolysis 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.3.1 The process of electrolysis 4.3.2 Electrolysis of molten ionic compounds 4.4.3.3 Using electrolysis to extract metals
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 4.4.3 Electrolysis 	 4.3.2.1 Moles (HT only) 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.5 Concentration of solutions 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.2.4 Limiting reactants of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction in terms of electrons (HT only) 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.3 Soluble salts 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.3.1 The process of electrolysis 4.3.2 Electrolysis of molten ionic compounds 4.3.3 Using electrolysis of aqueous solutions
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 4.4.3 Electrolysis 	 4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.2.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.4.2.1 Reactions of acids with metals 4.4.2.3 Soluble salts 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.3.1 The process of electrolysis 4.4.3 Electrolysis of molten ionic compounds 4.4.3 Electrolysis of aqueous solutions 4.4.3 Electrolysis of aqueous solutions 4.4.3 Ferentation of reactions at electrodes
4.3 Quantitative chemistry	4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 4.4.3 Electrolysis	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.1.4 Oxidation and reduction in terms of electrons (HT only) 4.4.2.1 Reactions of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.3 Soluble salts 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.3.1 The process of electrolysis 4.4.3.2 Electrolysis of molten ionic compounds 4.4.3.2 Electrolysis of aqueous solutions 4.4.3.4 Electrolysis of aqueous solutions 4.4.3.5 Representation of reactions at electrodes as half equations (HT only)
4.3 Quantitative chemistry	 4.3.2 Use of amount of substance in relation to masses of pure substances 4.3.3 Yield and atom economy of chemical reactions (chemistry only) 4.4.1 Reactivity of metals 4.4.2 Reactions of acids 4.4.3 Electrolysis 	4.3.2.1 Moles (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.2 Amounts of substances in equations (HT only) 4.3.2.3 Using moles to balance equations (HT only) 4.3.2.4 Limiting reactants (HT only) 4.3.2.5 Concentration of solutions 4.3.3.1 Percentage yield 4.3.3.2 Atom economy 4.3.4 Using concentrations of solutions in mol/dm3 (chemistry only) (HT only) 4.3.5 Use of amount of substance in relation to volumes of gases (chemistry only) (HT only) 4.1.1 Metal oxides 4.4.1.2 The reactivity series 4.4.1.3 Extraction of metals and reduction 4.4.2 Neutralisation of acids with metals 4.4.2.2 Neutralisation of acids and salt production 4.4.2.3 Soluble salts 4.4.2.4 The pH scale and neutralisation 4.4.2.5 Titrations (chemistry only) 4.4.3.1 The process of electrolysis 4.4.3.2 Electrolysis of molten ionic compounds 4.4.3.4 Electrolysis of aqueous solutions 4.4.3 Lectrolysis of aqueous solutions 4.4.3.4 Electrolysis of aqueous solutions

4.5 Energy changes		4.5.1.2 Reaction pro les
	4.5.1 Exothermic and endothermic reactions	4.5.1.3 The energy change of reactions (HT of
		4.5.2.1 Cells and batteries
	4.5.2 Chemical cells and fuel cells (chemistry only)	4.5.2.2 Fuel cells
		4.6.1.1 Calculating rates of reactions
		reactions
		4.6.1.3 Collision theory and activation energy
	4.6.1 Rate of reaction	4.6.1.4 Catalysts
		4.6.2.1 Reversible reactions
		4.6.2.2 Energy changes and reversible reaction
4.6 The rate and extent of chemical change		4.6.2.3 Equilibrium
		4.6.2.4 The effect of changing conditions on
	4.6.2 Reversible reactions and dynamic equilibrium	equilibrium (HT only)
		4.6.2.5 The effect of changing concentration
		4.6.2.6 The effect of temperature changes on
		equilibrium (HT only)
		4.6.2.7 The effect of pressure changes on
		equilibrium (HT only)
		4.7.1.1 Crude oil, hydrocarbons and alkanes
	4.7.1 Carbon compounds as fuels and	4.7.1.2 Fractional distillation and petrochem
	feedstock	4.7.1.3 Properties of hydrocarbons
		4.7.1.4 Cracking and alkenes
	4.7.2 Reactions of alkenes and alcohols	4.7.2.1 Structure and formulae of alkenes
4.7 Organic chemistry	(chemistry only)	4.7.2.2 Reactions of alkenes
		4.7.2.4 Carboxylic acids
		4.7.3.1 Addition polymerisation
		4.7.3.2 Condensation polymerisation (HT or
	4.7.3 Synthetic and naturally occurring	4.7.3.3 Amino acids (HT only)
	polymers (chemistry only)	4.7.3.4 DNA (deoxyribonucleic acid) and oth
		naturally occurring polymers
		4.8.1.1 Pure substances
	4.8.1 Purity, formulations and chromatography	4.8.1.2 Formulations
		4.8.1.3 Chromatography
	4.8.2 Identification of common gases	4.8.2.1 Test for hydrogen
		4.8.2.2 Test for oxygen
4.8 Chemical analysis		4.8.2.3 Test for carbon dioxide
		4.8.2.4 lest for chlorine
		4.8.3.2 Metal hydroxides
	4.8.3 Identification of ions by chemical and	4.8.3.3 Carbonates
	spectroscopic means (chemistry only)	4 8 3 4 Halides
		4 8 3 5 Sulfates
		4.8.3.6 Instrumental methods
		4.8.3.7 Flame emission spectroscopy
		4.9.1.1 The proportions of different gases in
		atmosphere
		4.9.1.2 The Earth's early atmosphere
	4.9.1 The composition and evolution of the Earth's	4.9.1.3 How oxygen increased
	atmosphere	4.9.1.4 How carbon dioxide decreased
4.9 Chemistry of the	4.9.2 Carbon dioxide and methane as	4.9.2.1 Greenhouse gases
autosphere	greenhouse gases	[4.9.2.2 Human activities which contribute to increase in greenhouse gases in the atmospheric
		4.9.2.3 Global climate change
		4.9.2.4 The carbon footprint and its reduction
		4.9.3.1 Atmospheric pollutants from fuels
	4.9.3 Common atmospheric pollutants and their	4.9.3.2 Properties and effects of atmospheric
	sources	pollutants
		4.10.1.1 Using the Earth's resources and
	4.10.1 Using the Earth's resources and obtaining potable water	sustainable development
		4.10.1.2 Potable water
		4.10.1.3 Waste water treatment
		4.10.1.4 Alternative methods of extracting m
		(HI only)
4.10 Using resources		(HI only) 4 10 2 1 Life cycle assessment
4.10 Using resources	4.10.2 Life cycle assessment and recycling	(<i>H1 only</i>) 4.10.2.1 Life cycle assessment 4.10.2.2 Ways of reducing the use of resource
4.10 Using resources	4.10.2 Life cycle assessment and recycling	(<i>H1 only</i>) 4.10.2.1 Life cycle assessment 4.10.2.2 Ways of reducing the use of resourc 4.10.3.1 Corrosion and its prevention
4.10 Using resources	4.10.2 Life cycle assessment and recycling	(H1 only) 4.10.2.1 Life cycle assessment 4.10.2.2 Ways of reducing the use of resource 4.10.3.1 Corrosion and its prevention 4.10.3.2 Alloys as useful materials
4.10 Using resources	4.10.2 Life cycle assessment and recycling 4.10.3 Using materials (chemistry only)	(H1 only) 4.10.2.1 Life cycle assessment 4.10.2.2 Ways of reducing the use of resourc 4.10.3.1 Corrosion and its prevention 4.10.3.2 Alloys as useful materials 4.10.3.3 Ceramics, polymers and composites
4.10 Using resources	4.10.2 Life cycle assessment and recycling 4.10.3 Using materials (chemistry only) 4.10.4 The Haber process and the use of NPK	(H1 only) 4.10.2.1 Life cycle assessment 4.10.2.2 Ways of reducing the use of resourc 4.10.3.1 Corrosion and its prevention 4.10.3.2 Alloys as useful materials 4.10.3.3 Ceramics, polymers and composites 4.10.4.1 The Haber process
4.10 Using resources	4.10.2 Life cycle assessment and recycling 4.10.3 Using materials (chemistry only) 4.10.4 The Haber process and the use of NPK fertilisers (chemistry only)	(H1 only) 4.10.2.1 Life cycle assessment 4.10.2.2 Ways of reducing the use of resource 4.10.3.1 Corrosion and its prevention 4.10.3.2 Alloys as useful materials 4.10.3.3 Ceramics, polymers and composite 4.10.4.1 The Haber process 4.10.4.2 Production and uses of NPK fertilis

Physics AQA Specification

Topic	Subtopic	Decription
•	•	4.1.1.1 Energy stores and systems
	4.1.1 En anna channaiste a suid-suid-suid-tha	4.1.1.2 Changes in energy
	4.1.1 Energy changes in a system, and the	4 1 1 3 Energy changes in systems
4.1 Energy	ways energy is stored before and after	4.1.1.7 Energy enanges in systems
	such changes	4.1.2.1 Energy transfers in a system
		4.1.2.1 Energy transfers in a system
	4.1.2 Conservation and dissipation of	4.1.2.2 Emclency
	energy	4.1.3 National and global energy resources
		4.2.1.1 Standard circuit diagram symbols
		4.2.1.2 Electrical charge and current
	4.2.1 Current, potential difference and	4.2.1.3 Current, resistance and potential difference
	resistance	4.2.1.4 Resistors
4.2 Electricity	4.2.2 Series and parallel circuits	
4.2 Liectricity		4.2.3.1 Direct and alternating potential difference
	4.2.3 Domestic uses and safety	4.2.3.2 Mains electricity
		4 2 4 1 Power
		4 2 4 2 Energy transfers in everyday appliances
	4.2.4 Enorgy transform	4.2.4.2 The National Crid
	4.2.4 Energy transfers	
		4.2.5.1 Static charge
	4.2.5 Static electricity (physics only)	4.2.5.2 Electric LEDS
	4.3.1 Changes of state and the particle	4.3.1.1 Density of materials
	model	4.3.1.2 Changes of state
		4.3.2.1 Internal energy
		4.3.2.2 Temperature changes in a system and speci c
4.3 Particle model of matter	4.3.2 Internal energy and energy	heat capacity
	transfers	4.3.2.3 Changes of heat and speci c latent heat
		4.3.3.1 Particle motion in gases
		4.3.3.2 Pressure in gases (physics only)
		4 3 3 3 Increasing the pressure of a gas (physics only)
	4.3.3 Particle model and pressure	(HT only)
		4 4 1 1 The structure of an atom
		4 4 1 2 Mass number atomic number and isotones
		4.4.1.2 Thus humber, atomic number and isotopes
	4.4.1 Atoms and isotones	(common content with chemistry)
	in in ritoriis unu isotopes	4 4 2 1 Radioactive decay and nuclear radiation
		4.4.2.2 Nuclear equations
		4.4.2.2 Nuclear equations
4.4 Atomic structure		4.4.2.3 Half-lives and the random nature of
	4.4.2 A tame and nuclear rediction	4.4.2.4 Padiagative contamination
	4.4.2 Atoms and nuclear radiation	
	4.4.3 Hazards and uses of radioactive	4.4.3.1 Background radiation
	emissions and of background radiation	4.4.3.2 Different half-lives of radioactive isotopes
	(physics only)	4.4.3.3 Uses of nuclear radiation
	4.4.4 Nuclear ssion and fusion (physics	4.4.4.1 Nuclear fission
	only)	4.4.4.2 Nuclear fusion
		4.5.1.1 Scalar and vector quantities
		4.5.1.2 Contact and non-contact forces
		4.5.1.3 Gravity
	4.5.1 Forces and their interactions	4.5.1.4 Resultant forces
	4.5.2 Work done and energy transfer	
	4.5.3 Forces and elasticity	
	4.5.4 Moments, levers and gears (physics	
4.5 Forces	only)	
	4.5.5 Pressure and pressure	4.5.5.1 Pressure in a fluid
	differences in uids (physics only)	4.5.5.2 Atmospheric pressure
		4 5 6 1 Describing motion along a line
		4.5.6.2 Forces accelerations and Newton's Laws of
		motion
	4.5.6 Forces and motion	4 5 6 3 Forces and braking
		4.5.7.1 Momentum is a property of moving objects
		4 5 7 2 Conservation of momentum
	157 Momentum (UT anhi)	4.5.7.2 Conservation of momentum
	4.5.7 Momentum (H1 Only)	4.5.7.5 Changes in momentum (physics only)
		4.0.1.1 Transverse and longitudinal waves
		4.6.1.2 Properties of waves
		4.6.1.3 Re ection of waves (physics only)
		4.6.1.4 Sound waves (physics only) (HT only)
		4.6.1.5 Waves for detection and exploration (physics
	4.6.1 Waves in air, uids and solids	only) (HT only)
4.6 Waves		4.6.2.1 Types of electromagnetic waves
		4.6.2.2 Properties of electromagnetic waves 1
		4.6.2.3 Properties of electromagnetic waves 2

Key Triple Science Only HL only

		4.6.2.4 Uses and applications of electromagnetic
		4.6.2.5 Lenses (physics only)
	4.6.2 Electromagnetic waves	4.6.2.6 Visible light (physics only)
		4.6.3.1 Emission and absorption of infrared radiation
	4.6.3 Black body radiation (physics only)	4.6.3.2 Perfect black bodies and radiation
	4.7.1 Permanent and induced	4.7.1.1 Poles of a magnet
	magnetism, magnetic forces and elds	4.6.3.2 Perfect black bodies and radiation
		4.7.2.1 Electromagnetism
		4.7.2.2 Fleming's left-hand rule (HT only)
4.7 Magnetism and		4.7.2.3 Electric motors (HT only)
electromagnetism	4.7.2 The motor effect	4.7.2.4 Loudspeakers (physics only) (HT only)
		4.7.3.1 Induced potential (HT only)
	473 Induced potential transformers and	4.7.3.2 Uses of the generator effect (HT only)
	the National Grid (physics only) (HT	4.7.3.3 Microphones (HT only)
	only)	4.7.3.4 Transformers (HT only)
		4.8.1.1 Our solar system
4.8 Space physics (physics		4.8.1.2 The life cycle of a star
only)	4.8.1 Solar system; stability of orbital	4.8.1.3 Orbital motion, natural and arti cial satellites
	motions; satellites (physics only)	4.8.2 Red-shift (physics only)